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Common compounds of boron include borax and boric acid (H₃BO₃). Boron is the only element in group 3 that is not a metal. It has properties that lie between metals and non-metals (semimetals). For example Boron is a semiconductor unlike the rest of the group 13 elements.

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Flame-retardant boron compounds such as boric acid (B₂O₃·3H₂O/B(OH)₃), boron oxide (B₂O₃), boron nitride, borosilicates, zinc borates, etc. (Fig. 2.18) are excellent char reinforcers. Some boron compounds—especially borosilicates—form glasses at higher temperatures, which act as a heat barrier and prevent the dripping of epoxy resins.

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Boron compounds include all chemical compounds including the element boron. For an alternative listing of boron compounds, please check inorganic_compounds_by_element#Boron. Wikimedia Commons has media related to Boron compounds.

Boron compounds play an important role in making glass, ceramics, bleach, semiconductors and magnets. There are two important compounds of Boron - Boric Acid (H₃BO₃) and Sodium tetraborate pentahydrate (Na₂B₄O₇·5H₂O). Most of sodium tetraborate pentahydrate are used in the manufacture of fiberglass insulation and bleach.

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Some compounds of boron. Some other important compounds of boron are borax, orthoboric acid and boron hydrides. Let us know something about these compounds. Borax: Borax is an important boron compound. It has a white crystalline solid structure with formula of Na₂B₄O₇·10H₂O. The actual formula for borax is Na₂[B₄O₅(OH)₄]. 8H₂O

Common Compounds Of Boron | Characteristics Of Chemical Boron

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In compounds of boron (where known), the most common oxidation numbers of boron are: 3. Hydrides. The term hydride is used to indicate compounds of the type M_xH_y and not necessarily to indicate that any compounds listed behave as hydrides chemically. Diborane (6): B₂H₆; Decaborane (14): B₁₀H₁₄; Hexaborane (10): B₆H₁₀; Pentaborane (9): B₅H₉

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Boron occurs combined as borax, kernite, and tincalconite (hydrated sodium borates), the major commercial boron minerals, especially concentrated in the arid regions of California, and as widely dispersed minerals such as colemanite, ulexite, and tourmaline. Sassolite—natural boric acid—occurs especially in Italy.

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The reaction of boron carbide with bromine at temperatures above 300 °C leads to the formation of boron tribromide. The product can be purified by vacuum distillation. History. The first synthesis was done by M. Poggiale in 1846 by reacting boron trioxide with carbon and bromine at high temperatures: B₂O₃ + 3 C + 3 Br₂ → 2 BBr₃ + 3 CO

Boron tribromide - Wikipedia

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Chemistry of Boron (Z=5) - Chemistry LibreTexts

Boron (B) is a chemical element with an atomic number 5 that belongs in the Period 2 and Group 13 in the periodic table. It is a low-abundant metalloid that is a poor electrical conductor at room temperature. Natural boron exists in 2 stable isotopes however pure boron is hard to prepare due to contamination by different elements. It is also found for applications including ceramics, high ...

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In this case, the anion is the so-called COSAN anion (its full name is cobaltabisdicarbollide, and chemical formula [3,3'-Co (C₂B₉H₁₁)₂]⁻), a boron cluster-based complex with a Co (III)...

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